

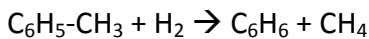
Process Systems Engineering

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Exercise #2

The kinetic model relative to the main reactions that schematize the process of dealkylation is shown below.

Reaction #1:

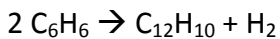


$$A_1 = 3.5\text{E} + 10$$
$$E_1 = 50900$$

$$k_1 = A_1 \exp\left(-\frac{E_1}{RT}\right)$$

$$R_1 = k_1 c_T \sqrt{c_H}$$

Reaction #2:



$$A_2 = 2.1\text{E} + 12$$
$$E_2 = 60500$$

$$k_2 = A_2 \exp\left(-\frac{E_2}{RT}\right)$$

$$R_2 = k_2 c_B^2$$

Activation energies are in kcal/kmol, and the rates of reaction are expressed in kmol/m³/s. We

ask to:

- determine, via numerical integration of the plug-flow model of the reactor, the conversion, the selectivity and the residence time as a function of the operating temperature, by assuming isothermal the reactor, and neglecting the presence of recycles in the evaluation of initial concentrations
- evaluate the adiabatic ΔT of the reaction so to determine if the reactor can be considered isothermal

Carry out the following diagrams:

- conversion/selectivity as a function of reactor temperature
- temperature/conversion by imposing the selectivity of at least 96%
- temperature/residence time of imposing the selectivity of at least 96%
- residence time/flow rates of each of the input streams as a function of the reactor temperature

A suitable range of investigation for the working temperature of this unit is: 600-750 °C.